

AR Chemistry Notes
Equilibrium 2

Name _____

Value of Keq:

Large Number (eg. 200000000) _____

Medium Number (eg. 100) _____

Small Number (eg. 0.00000006) _____

LeChatelier's Principle

Definition: _____

Why PARTIALLY?

Ex: $2 \rightleftharpoons 2$

Add 4 

How Do The Reaction Shift?

STRESS

COUNTER – SHIFT

Add Reactants _____

Add Products _____

Increase Temperature (Add Heat) _____

Increase Pressure (Decrease Volume) _____

Remove Reactants _____

Remove Products _____

Lower Temperature (Remove Heat) _____

Lower Pressure (Increase Volume) _____

STRESS**COUNTER – SHIFT**

Add Reactants

Shifts →

Add Products

Shifts ←

Increase Temperature (Add Heat [+])

Shifts to side that removes heat (ENDO +)

Increase Pressure (Decrease Volume)

Shifts to side that has **fewer** moles **gas**

Remove Reactants

Shifts ←

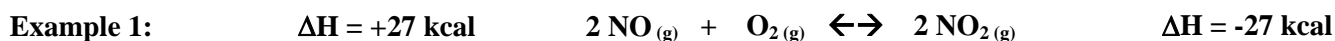
Remove Products

Shifts →

Lower Temperature (Remove Heat [-])

Shifts to side that releases heat (EXO -)

Lower Pressure (Increase Volume)

Shifts to side that has **more** moles **gas**

Which way will it shift?

Add Reactants

Add Products

Increase Temp : Add Heat

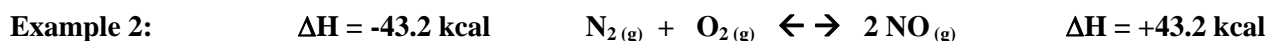
Increase Pressure: Decrease Volume

Remove Reactants

Remove Products

Lower Temp : Remove Heat

Lower Pressure : Increase Volume



Which way will it shift?

Add Reactants

Add Products

Increase Temp : Add Heat

Increase Pressure: Decrease Volume

Remove Reactants

Remove Products

Lower Temp : Remove Heat

Lower Pressure : Increase Volume



Which way will it shift?

Add Reactants

Lower Temp _____

Add Products

Lower Pressure _____

Increase Temp : Add Heat

Increase Pressure: Decrease Volume

Remove Reactants

Remove Products
