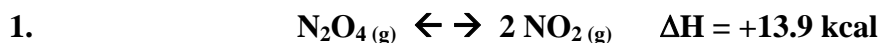
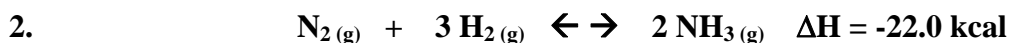


AR Chemistry: LeChatelier's Principle Review Set



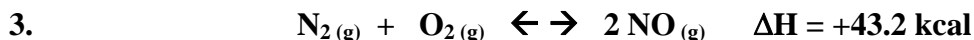
Which way will a chemical reaction shift when the following stresses are put on it:

- | | |
|-------------------------------------|-------------------------------------|
| a. Reactants are added | b. Reactants are removed |
| c. Products are added | d. Products are removed |
| e. Temperature is increased | f. Temperature is decreased |
| g. Pressure is increased | h. Pressure is decreased |
| i. Volume of container is increased | j. Volume of container is decreased |



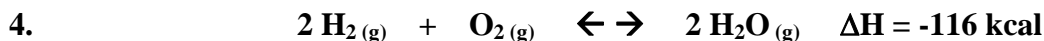
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What do you do the system to shift the equilibrium to the creation of more products (increase the yield)?

- Reactants
- Products
- Temperature
- Pressure
- Size of Container

Answers

Change to System	Problem 1	Problem 2	Problem 3	Problem 4	Problem 5
Add Reactants	Right	Right	Right	Right	Right
Remove Reactants	Left	Left	Left	Left	
Add Products	Left	Left	Left	Left	
Remove Products	Right	Right	Right	Right	Right
Increase Temperature	Endo Right	Endo Left	Endo Right	Endo Left	Endo Right
Decrease Temperature	Exo Left	Exo Right	Exo Left	Exo Right	
Increase Pressure	Fewer mol gas Left	Fewer mol gas Right	Fewer mol gas No Shift	Fewer mol gas Right	
Decrease Pressure	More mol gas Right	More mol gas Left	More mol gas No Shift	More mol gas Left	More mol gas Right
Increase Volume	More mol gas Right	More mol gas Left	More mol gas No Shift	More mol gas Left	More mol gas Right
Decrease Volume	Fewer mol gas Left	Fewer mol gas Right	Fewer mol gas No Shift	Fewer mol gas Right	