

## 1. Set It Up

1 Mole = 6.02 E23 atoms/molecules = Molec. Mass in grams = 22.4 L @ STP

How many grams are in 2.50 moles of CO?

2.50 moles CO  $\xrightarrow{\hspace{10em}}$  g CO

## 2. Move the Words

1 Mole = 6.02 E23 atoms/molecules = Molec. Mass in grams = 22.4 L @ STP

2.50 moles CO  $\xrightarrow{\hspace{10em}}$  moles CO  $\xleftarrow{\hspace{10em}}$  g CO

## 3. Put in Numbers

1 Mole = 6.02 E23 atoms/molecules = Molec. Mass in grams = 22.4 L @ STP

2.50 moles CO  $\xrightarrow{\hspace{10em}}$  1 moles CO  $\xleftarrow{\hspace{10em}}$  28.01 g CO  $\xleftarrow{\hspace{10em}}$  g CO

C: 1 x 12.01 = 12.01 O: 1 x 16.00 = 16.00 = 28.01
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## 4. Do The Math

1 Mole = 6.02 E23 atoms/molecules = Molec. Mass in grams = 22.4 L @ STP

~~2.50 moles CO~~ x  $\frac{28.01 \text{ g CO}}{\cancel{1 \text{ moles CO}}}$  = 77.0 g CO

