

## A. Gram to Gram

11. $3 \text{H}_2\text{SO}_4 + 2 \text{Al}(\text{OH})_3 \rightarrow \text{Al}_2(\text{SO}_4)_3 + 6 \text{H}_2\text{O}$	Grams $\text{H}_2\text{O}$ from 98.0 g $\text{Al}(\text{OH})_3$ ? (67.8)
12. $2 \text{C}_6\text{H}_6 + 15 \text{O}_2 \rightarrow 12 \text{CO}_2 + 6 \text{H}_2\text{O}$	How many grams of water form when 34.5 g $\text{C}_6\text{H}_6$ react? (23.9)
13. $2 \text{CH}_3\text{OH} + 3 \text{O}_2 \rightarrow 2 \text{CO}_2 + 4 \text{H}_2\text{O}$	How many grams of $\text{CO}_2$ form if 90.0 g water form? (110)
14. $3 \text{Fe} + 2 \text{O}_2 \rightarrow \text{Fe}_3\text{O}_4$	How many grams of $\text{Fe}_3\text{O}_4$ form when 78.0 g Fe react? (108)
15. $2 \text{Pb}(\text{NO}_3)_2 \rightarrow 2 \text{PbO} + 4 \text{NO}_2 + \text{O}_2$	How many grams of $\text{NO}_2$ form if 129 g $\text{PbO}$ form? (53.2)

## B. Do 5-10 Balancing Equations from your Problem Sets